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(54) Title: NON-AZEOTROPIC REFRIGERANT COMPOSITIONS COMPRISING DIFLUOROMETHANE; 1,1,1-TRIFLUOROETHANE; OR PROPANE			
(57) Abstract  The present invention provides refrigerant blends which are replacements for chlorodifluoromethane (HCFC-22). The present blends have refrigeration characteristics which are similar to HCFC-22. The blends comprise from about 10 to about 90 weight percent of a first component selected from the group consisting of 1,1,1-trifluoroethane, difluoromethane, propane, and mixture thereof; from about 1 to about 50 weight percent of a second component selected from the group consisting of hydrofluorocarbon having 1 to 3 carbon atoms, fluorocarbon having 1 to 3 carbon atoms, inorganic compound, and mixtures thereof having a boiling point at atmospheric pressure in the range from about -90 degrees C to less than -50 degrees C; and from about 1 to about 50 weight percent of a third component which is hydrofluorocarbon having 1 to 3 carbon atoms, other than 1,1,1-trifluoroethane, having a boiling point at atmospheric pressure in the range from about -50 degrees C to about -10 degrees C. The refrigerant compositions have a vapor pressure substantially equal to the vapor pressure of HCFC-22.			

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NON-AZEOTROPIC REFRIGERANT COMPOSITIONS COMPRISING  
DIFLUOROMETHANE; 1,1,1-TRIFLUOROETHANE; OR PROPANE

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FIELD OF THE INVENTION

This invention relates to novel nonazeotropic compositions containing difluoromethane; 1,1,1-trifluoroethane; or propane. These mixtures have improved efficiency and capacity as refrigerants for heating and cooling.

BACKGROUND OF THE INVENTION

15 Fluorocarbon based fluids have found widespread use in industry for refrigeration, air conditioning and heat pump applications. Vapor compressions cycles are one form of refrigeration. In its simplest form, the vapor compression cycle involves changing the refrigerant from the liquid to the vapor phase through heat absorption at a low pressure, and then from the vapor to the liquid phase through heat removal at an elevated pressure. First, the refrigerant is vaporized in the evaporator which is in contact with the body to be cooled. The pressure in the evaporator is such that the boiling point of the refrigerant is below the temperature of the body to be cooled. Thus, 20 heat flows from the body to the refrigerant and causes the refrigerant to vaporize. The formed vapor is then removed by means of a compressor in order to maintain the low pressure in the evaporator. The temperature and pressure of the vapor are then raised through the addition of mechanical energy by the compressor. The high pressure vapor then passes to the condenser whereupon heat exchange with a cooler medium, the sensible and latent heats are removed with subsequent condensation. The hot liquid refrigerant then passes to the expansion valve and is ready to cycle again.

35 While the primary purpose of refrigeration is to remove energy at low temperature, the primary purpose of a heat pump is to add energy at higher temperature. Heat pumps are considered reverse cycle systems



be redesigned, see for example T. Atwood, "NARBS-The Promise and the Problems", paper 86-WA/HT-61 American Society of Mechanical Engineers. In most existing designs of refrigeration equipment, a temperature glide is a cause of concern. Therefore, nonazeotropic refrigerant mixtures have not found wide use. An environmentally acceptable nonazeotropic mixtures with a small temperature glide and with an advantage in refrigeration capacity over other known pure fluids will have a general commercial interest.

Chlorodifluoromethane (HCFC-22) is a currently used refrigerant. Although HCFC-22 is only partially halogenated, it still contains chlorine and hence has a propensity for ozone depletion. What is needed in the refrigerant art is a replacement for HCFC-22 which has similar refrigeration characteristics, is nonflammable, has low temperature guides, and contains no ozone-depleting chlorine atoms.

U.S. Patent 4,810,403 teaches ternary or higher blends of halocarbon refrigerants which are substitutes for dichlorodifluoromethane (CFC-12). The blends have a first component which has a boiling point at atmospheric pressure in the range of -50 degrees C to -30 degrees C, a second component which has a boiling point at atmospheric pressure in the range of -30 degrees C to -5 degrees C, and a third component which has a boiling point at atmospheric pressure in the range of -15 degrees C to 30 degrees C. The preferred blend contains chlorodifluoromethane (HCFC-22), 1,1-difluoroethane (HFC-152a), and 1,2-dichloro-1,1,2,2-tetrafluoroethane (CFC-114). As the reference lists HCFC-22 as a possible refrigerant component, the reference is not teaching refrigerant substitutes for HCFC-22.

As such, the art is seeking new fluorocarbon based mixtures which offer alternatives for HCFC-22 in refrigeration and heat pump applications. Currently, of particular interest, are fluorocarbon based mixtures which are considered to be environmentally acceptable substitutes for the presently used hydrochlorofluorocarbons which are

suspected of causing environmental problems in connection with the earth's protective ozone layer. Mathematical models have substantiated that hydrofluorocarbons, such as 1,1,1-trifluoroethane (HFC-143a) or difluoromethane (HFC-32) will not adversely affect atmospheric chemistry, being negligible contributors to stratospheric ozone depletion and global warming.

The substitute materials must also possess those properties unique to the CFC's including chemical stability, low toxicity, non-flammability, and efficiency in-use. The latter characteristic is important, for example, in air conditioning and refrigeration where a loss in refrigerant thermodynamic performance or energy efficiency may have secondary environmental impacts through increased fossil fuel usage arising from an increased demand for electrical energy.

The aforementioned environmentally acceptable refrigerants HFC-32 and HFC-143a are flammable which may limit their general use. These refrigerants are generally regarded as too low boiling fluids to directly replace chlorodifluoromethane (HCFC-22).

In order to overcome the flammability of HFC-32, we blended HFC-32 with 1,1,1,2-tetrafluoroethane (HFC-134a) and the result was zero ozone depletion potential compositions which are useful substitutes for HCFC-22. At high amounts of HFC-32 though, compositions of HFC-32 and HFC-134a are flammable. In order to completely eliminate the flammability of such compositions, we decided to add a third nonflammable component. In adding a third component, we wanted the resulting ternary composition to have a zero ozone depletion potential and have a boiling point comparable to that of HCFC-22. One member from the list of compounds having zero ozone depletion potential and boiling points at atmospheric pressure in the range of -90 degrees C to -60 degrees C is trifluoromethane (HFC-23) which has a low critical temperature; as those skilled in the art know, compounds having low critical temperatures are not used as refrigerants because they do not condense at room temperature and in a refrigerant blend,

would be expected to substantially reduce the refrigeration efficiency and capacity of the blend. We were pleasantly surprised to find that in addition to being nonflammable, a blend of HFC-32, HFC-134a, and HFC-23 has refrigeration efficiency and capacity substantially the same as a  
5 blend of HFC-32 and HFC-134a.

#### SUMMARY OF THE INVENTION

Thus, we have discovered refrigerant blends which  
10 are substitutes for HCFC-22. These nonazeotropic refrigerant compositions comprise from about 10 to about 90 weight percent of a first component selected from the group consisting of 1,1,1-trifluoroethane (HFC-143a), difluoromethane (HFC-32), propane, and mixtures thereof; from about 1  
15 to about 50 weight percent of a second component selected from the group consisting of hydrofluorocarbon having 1 to 3 carbon atoms, fluorocarbon having 1 to 3 carbon atoms, inorganic compound, and mixtures thereof having a boiling point at atmospheric pressure in the range from about -90 degrees C to less than -50 degrees C; and from  
20 about 1 to about 50 weight percent of a third component which is hydrofluorocarbon having 1 to 3 carbon atoms, other than 1,1,1-trifluoroethane, having a boiling point at atmospheric pressure in the range from about -50 degrees C to about -10 degrees C. The refrigerant compositions have a vapor pressure substantially equal to the vapor  
25 pressure of HCFC-22.

The term "hydrofluorocarbon" as used herein means a compound having carbon, hydrogen, and fluorine atoms. The term "fluorocarbon" as used herein means a compound having carbon and fluorine atoms. For the  
30 second component, any hydrofluorocarbon having 1 to 3 carbon atoms, fluorocarbon having 1 to 3 carbon atoms, or inorganic compound having a boiling point at atmospheric pressure in the range from about -90 degrees C to less than -50 degrees C may be used in the present invention. For the third component, any hydrofluorocarbon having 1 to 3  
35 carbon atoms, other than 1,1,1-trifluoroethane, having a boiling point

at atmospheric pressure in the range from about -50 degrees C to about -10 degrees C may be used in the present invention.

The preferred first component is difluoromethane.

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Preferably, the second component is selected from the group consisting of: trifluoromethane (HFC-23), hexafluoroethane (FC-116), carbon dioxide or sulphur hexafluoride. The preferred second component is trifluoromethane. All members listed for the second  
10 component are nonflammable and generally boil at a temperature below that of HFC-32 or HFC-143a.

Preferably, the third component is selected from the group consisting of: pentafluoroethane (HFC-125), 1,1,2,2-tetrafluoroethane  
15 (HFC-134), 1,1,1,2-tetrafluoroethane (HFC-134a), 1,1,1,2,3,3,3-heptafluoropropane (HFC-227ea), 1,1,1,2,2,3,3-heptafluoropropane (HFC-227ca), or 1,1,1,2,2-pentafluoropropane (HFC-245cb). The preferred third component is 1,1,1,2-tetrafluoroethane. All members listed for the third component are nonflammable and generally boil at a  
20 temperature above that of HFC-32 or HFC-143a.

Small quantities of HFC-227ea, HFC-227ca, and HFC-245Cb are available from PCR and Halocarbon Products. All other components of the present invention are available in commercial quantities. Also, HFC-  
25 227ea, HFC-227ca, and HFC-245cb may be prepared according to known methods such as those disclosed in International Publication Number WO 90/08754. For example, HFC-227ca may be prepared by reacting 1,1,1,3,3-pentachloro-2,2-difluoropropane with niobium pentachloride at 120 degrees C. HFC-245cb may be prepared by reacting 1,1,1,2,2-  
30 pentachloropropane with tantalum pentafluoride at 120 degrees C.

By "vapor pressure substantially equal to the vapor pressure of chlorodifluoromethane" or "similar refrigeration characteristics" is meant a vapor pressure which is plus or minus 30  
35 percent of the vapor pressure of HCFC-22 at the same temperature over



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the temperature range of about 0 degrees C to about 100 degrees C.

Additional components may be added to the mixture to tailor the properties of the mixtures according to the need.

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Other advantages of the invention will become apparent from the following description.

#### DESCRIPTION OF THE DRAWING

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The Figure illustrates nonflammable compositions of the present invention.

#### DETAILED DESCRIPTION OF THE INVENTION

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The properties of the preferred components of the present invention are listed in Table 1 below. BP in Table 1 stands for Boiling Point while CT stands for Critical Temperature. The \* in Table 1 means sublimates at one atm pressure and the boiling point is the triple point.

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TABLE 1

	<u>No.</u>	<u>Formula</u>	<u>BP(degrees C)</u>	<u>CT(degrees C)</u>
5	HFC-32	CH <sub>2</sub> F <sub>2</sub>	-51.7	78.4
	HFC-143a	CF <sub>3</sub> CH <sub>3</sub>	-47.6	73.1
	Propane	C <sub>3</sub> H <sub>8</sub>	-42.1	96.7
	HFC-23	CHF <sub>3</sub>	-82.1	25.9
	FC-116	C <sub>2</sub> F <sub>6</sub>	-78.1	24.3
10	---	CO <sub>2</sub> *	-78.5	31.3
	---	SF <sub>6</sub>	-64.0	45.5
	HFC-125	C <sub>2</sub> HF <sub>5</sub>	-48.5	66.3
	HFC-134	CHF <sub>2</sub> CHF <sub>2</sub>	-19.7	118.9
	HFC-134a	CF <sub>3</sub> CH <sub>2</sub> F	-26.5	101.1
15	HFC-227ea	CF <sub>3</sub> CFHCF <sub>3</sub>	-16.5	102.0
	HFC-227ca	CF <sub>3</sub> CF <sub>2</sub> CHF <sub>2</sub>	-15.6	104.7
	HFC-245cb	CF <sub>3</sub> CF <sub>2</sub> CH <sub>3</sub>	-17.5	107.0

The most preferred composition comprises difluoromethane,  
 20 trifluoromethane, and 1,1,1,2-tetrafluoroethane.

In a preferred embodiment of the invention, the compositions  
 comprise from about 20 to about 80 weight percent of the first  
 component, from about 2 to about 40 weight percent of the second  
 25 component, and from about 2 to about 40 weight percent of the third  
 component.

In one process embodiment of the invention, the compositions of  
 the invention may be used in a method for producing refrigeration which  
 30 involves condensing a refrigerant comprising the compositions and  
 thereafter evaporating the refrigerant in the vicinity of the body to  
 be cooled.

In another process embodiment of the invention, the compositions of the invention may be used in a method for producing heating which involves condensing a refrigerant comprising the compositions in the vicinity of the body to be heated and thereafter  
 5 evaporating the refrigerant.

Preferably the components used should be of sufficiently high purity so as to avoid the introduction of adverse influences upon the properties of the system.  
 10

As mentioned above, when a refrigerant composition contains a flammable component like HFC-32, HFC-143a, or propane, the possibility of either the leaking vapor or the remaining liquid becoming flammable is a very undesirable hazard. We have discovered  
 15 that the claimed compositions of the refrigerant blends containing either HFC-32, HFC-143a, or propane can be so formulated with the components from the two nonflammable groups that the original composition is nonflammable and the leaking vapor as well as the remaining liquid never become flammable.  
 20

The present invention comprises ternary and higher blends based either on HFC-32, HFC-143a, or propane that have a vapor pressure substantially the same as the vapor pressure of HCFC-22 and which retain this relationship even after substantial evaporation losses,  
 25 e.g. up to 50 percent by weight. A vapor pressure temperature relationship similar to HCFC-22 is especially desirable because it will need minimum amount of modifications in the present refrigeration equipment which is designed around the vapor pressure temperature relationship of the HCFC-22.  
 30

It should be understood that the present compositions may include additional components so as to form new compositions. Any such compositions are considered to be within the scope of the present invention as long as the compositions have essentially the same  
 35 characteristics and contain all the essential components described

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herein.

The present invention is more fully illustrated by the following non-limiting Examples.

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EXAMPLE 1

This Example shows that it is possible to calculate the thermodynamic properties of a ternary mixtures from using equation of state techniques. These are important for estimating theoretical performance of a refrigerant as discussed in Example 3. The equation of state package used was based on the NIST Mixtures Properties formalism (DDMIX) available from the National Institute of Standards and Technology, Gaithersberg, Maryland 20899. An example of measured and calculated bubble pressures of a 48.1 wt% HFC-23, 19.3 wt% HFC-32, and 32.6 wt% HFC-134a ternary nonazeotropic blend is shown in Table 2. The very good agreement shows the high degree of confidence that may be placed in the results of the experiments and the theory.

20

TABLE 2

	<u>Temperature/K</u>	<u>Bubble Pressure exptl., psia</u>	<u>Bubble Pressure calcd., psia</u>
25	263.54	154.2	151.8
	268.49	176.4	174.8
	278.38	230.0	228.2
	288.09	293.4	290.9
	298.08	367.9	366.9
30	308.09	453.4	455.3
	318.12	550.7	556.1

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EXAMPLE 2

By preparing various compositions of HFC-134a/HFC-32/HFC-23 in air and determining their flammability, it is possible to map out the region of compositions in air that are flammable. See, e.g. P.A. Sanders, The Handbook of Aerosol Technology at 146 (2d Ed. 1979). The maximum amount of HFC-32 that can be blended with HFC-134a and HFC-23 and remain nonflammable in all proportions in air, can be determined from such a plot. Table 3 summarizes the maximum or critical composition of HFC-32 attainable with HFC-134a and a higher pressure component (e.g. HFC-23, HF-116, SF<sub>6</sub>, and CO<sub>2</sub>) for the binary mixtures. The CFR is the critical flammability ratio: which is the maximum amount of HFC-32 that a mixtures of HFC-32/X can contain and still be nonflammable in all proportions in air. X represents the higher pressure components listed in Table 3. These binary flammability data can be used to predict the flammability of the more complex ternary mixture plus air. This complex mixture of three components and air does not lend itself to simple ternary diagrams. Therefore, air is not included so that we are able to show the data graphically. The air proportion itself is not important just whether or not the mixture is flammable in some proportion with air. Figure 1 shows a composition of HCFC-134a, HFC-32, and HFC-23. Above the line A-B (more HFC-32), mixtures of those compositions are flammable in some proportion in air while below line A-B (less HFC-32), mixtures of those compositions are not flammable in air at any proportion of air. Further this diagram depicts compositions that will remain nonflammable in the event of a vapor leak. If the leak is from the liquid phase, some liquid will vaporize to fill the space vacated by the leaking liquid. Because the vapor is 1/25th as dense as the liquid, and very little vaporization occurs, therefore, very little fractionation occurs. In contrast, when the vapor phase is removed, all the liquid is eventually vaporized, producing a dramatic amount of fractionation. Liquid leaks produce only minor changes in the composition of the mixture. As such, a liquid leak is not problematic and only the case of a vapor leak must be considered.

Shifts in the compositions of the vapor and liquid phases during leaking were calculated using ideal solution behavior. These types of calculations were used to determine what starting compositions would remain nonflammable on evaporation. Line D-C in Figure 1 separates those compositions that could have flammable liquid phase compositions from those compositions that would remain nonflammable. Compositions rich in HFC-134a (right of the line) would have liquid phase compositions that remain nonflammable on evaporation. Line C-E separates composition that would fractionate given vapors that are flammable from those that would not produce flammable vapors. Compositions having more, HFC-23 (left of the line) would remain nonflammable vapors on segregation. Therefore, compositions below line D-C-E would not fractionate to produce either liquid or vapor phases that could be flammable.

TABLE 3

<u>Gas in HFC-32</u>	<u>Maximum HFC-32 Compo (mole or volume %)</u>	<u>% air at CFR (mole or volume %)</u>
HFC-134a	72.9	20
HFC-23	75.3	19
HFC-116	88.1	20
SF <sub>6</sub>	87.9	21
CO <sub>2</sub>	55.2	29

EXAMPLE 3

This Example shows that a HFC-32 containing blend has a performance similar to HCFC-22, yet is nonflammable even after substantial vapor leakage.

The theoretical performance of a refrigerant at specific operating conditions can be estimated from the thermodynamic properties of the refrigerant using standard refrigeration cycle analysis techniques, see for example, "Fluorocarbons Refrigerants Handbook". Ch.

3, Prentice-Hall, (1988), by R.C. Downing. The coefficient of performance, COP, is a universally accepted measure, especially useful in representing the relative thermodynamic efficiency of a refrigerant in a specific heating or cooling cycle involving evaporation or  
 5 condensation of refrigerant. In refrigeration engineering, this term expresses the ratio of useful refrigeration to the energy applied by the compressor in compressing the vapor. The capacity of a refrigerant represents the volumetric efficiency of the refrigerant. To a compressor engineer, this value expresses the capability of a  
 10 compressor to pump quantities of heat for a given volumetric flow rate of refrigerant. In other words, given a specific compressor, a refrigerant with a higher capacity will deliver more cooling or heating power. A similar calculation can also be performed for nonazeotropic refrigerant blends.

15 We have performed this type of calculation for packaged air conditioning cycle where the condenser temperature is typically 115°F and the evaporator temperature is typically 40°F. We have further assumed isentropic compression and a compressor inlet temperature of  
 20 60°F. Such calculations were performed for a 0.72/28.71/70.57 by weight blend of HFC-23, HFC-32, and HFC-134a. The temperature glide in typical HCFC-22 application in no case exceeded 15°F. The coefficient of performance (COP), a measure of energy efficiency of the fluid, was found to be 5.36 as compared to 5.51 found for HCFC-22 in the same  
 25 conditions. According to the known art (D.A. Didion and D.M. Livens "The Role of Refrigerant Mixtures as Alternatives" in CFC's: Today's Options..... Tomorrow's Solutions, NIT, 1990), the temperature glides of the order of 10°F are minor enough for the mixture to be termed Near-Azeotropic. Therefore, the temperature glide of the mixture  
 30 composition claimed is small enough and does not pose a problem for conventional refrigeration units. As can be seen from the attached Figure 1, which gives the flammability limits of the three component blend of HFC-23, HFC-32, and HFC-134a measured by an ASTM 681 apparatus, the blend is nonflammable. Its vapor pressure is 11.37 bars  
 35 at 25°C within 10 percent of the HCFC-22 vapor pressure. The

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refrigeration capacity is about 95% of the HCFC-22. After 50 weight percent of the refrigerant is lost through the leakage of the vapor, the vapor pressure of the blend is 9.44 bars, still within 10% of the HCFC-22 value. The refrigeration capacity has decreased to only 83% of the HCFC-22 value. The COP of the remaining fluid remained substantially the same at 5.37. Both the vapor at 46 volume percent HFC-32 and the liquid at 28 volume percent HFC-32 has remained nonflammable as seen from Figure 1.

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#### EXAMPLE 4

We have performed another calculation of the type given in Example 3 for packaged air conditioning cycle where the condenser temperature is typically 115°F and the evaporator temperature is typically 40°F. We have further assumed isentropic compression and a compressor inlet temperature of 60°F. This time such calculations were performed for a 77.56 gram blend of 0.0384 moles of HFC-23, 0.4648 moles of HFC-32, and 0.4968 moles of HFC-134a. The temperature glide in typical HCFC-22 application in no case exceeded 17°F. As can be seen from the attached Figure 1, which gives the flammability limits of the three component blend of HFC-23, HFC-32, and HFC-134a measured by an ASTM 681 apparatus, the blend is nonflammable. Its vapor pressure is 12.43 bars at 25°C within 25 percent of the HCFC-22 vapor pressure. The refrigeration capacity is substantially the same as the HCFC-22. The COP was 5.13. After 50 weight percent of the refrigerant is lost through the leakage of the vapor, the vapor pressure of the blend is 10.08 bars, within 2% of the HCFC-22 value. The refrigeration capacity has decreased to only 87% of the HCFC-22 value. The COP has increased marginally to 5.35. Both the vapor at 51 volume percent HFC-32 and the liquid at 33 volume percent HFC-32 has remained nonflammable as seen from Figure 1.

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EXAMPLE 5

We have performed another calculation of the type given in Examples 3 and 4 under the conditions given earlier. This time such calculations were performed for a 75.62 gram blend of 0.0651 moles of HFC-23, 0.4865 moles of HFC-32, and 0.4484 moles of HFC-134a. The temperature glide in typical HCFC-22 application in no case exceeded 20°F. As can be seen from the attached Figure 1, which gives the flammability limits of the three component blend of HFC-23, HFC-32, and HFC-134a measured by an ASTM 681 apparatus, the blend is nonflammable. Its vapor pressure is 13.38 bars at 25°C within 30 percent of the HCFC-22 vapor pressure. The refrigeration capacity is substantially the same as the HCFC-22. The COP is 5.02. After 50 weight percent of the refrigerant is lost through the leakage of the vapor, the vapor pressure of the blend is 10.78 bars, within 4% of the HCFC-22 value. The refrigeration capacity has decreased to only 91% of the HCFC-22 value. The COP is now 5.31. Both the vapor at 54 volume percent HFC-32 and the liquid at 37 volume percent HFC-32 has remained nonflammable as seen from Figure 1.

20

Having described the invention in detail and by reference to preferred embodiments thereof, it will be apparent that modifications and variations are possible without departing from the scope of the invention defined by the claims.

25

What is claimed is

1. Refrigerant compositions comprising from about  
10 to about 90 weight percent of a first component selected from the  
5 group consisting of 1,1,1-trifluoroethane, difluoromethane, propane,  
and mixtures thereof; from about 1 to about 50 weight percent of a  
second component selected from the group consisting of  
hydrofluorocarbon having 1 to 3 carbon atoms, fluorocarbon having 1 to  
3 carbon atoms, inorganic compound, and mixtures thereof having a  
10 boiling point at atmospheric pressure in the range from about -90  
degrees C to less than -50 degrees C; and from about 1 to about 50  
weight percent of a third component which is hydrofluorocarbon having 1  
to 3 carbon atoms, other than 1,1,1-trifluoroethane, having a boiling  
point at atmospheric pressure in the range from about -50 degrees C to  
15 about -10 degrees C wherein said refrigerant compositions have a vapor  
pressure substantially equal to the vapor pressure of  
chlorodifluoromethane.

2. The refrigerant compositions of claim 1 wherein said first,  
20 second, and third components and their weight percents are selected so  
that the resulting refrigerant compositions are nonflammable.

3. The refrigerant compositions of claim 1 comprising from  
about 20 to about 80 weight percent said first component, from about 2  
25 to about 40 weight percent said second component, and from about 2 to  
about 40 weight percent said third component.

4. The refrigerant compositions of claim 1 wherein said first  
component is 1,1,1-trifluoroethane.

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5. The refrigerant compositions of claim 4 wherein  
said second component is selected from the group  
consisting of trifluoromethane, hexafluoroethane, carbon  
dioxide, and sulfur hexafluoride.

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6. The refrigerant compositions of claim 4 wherein said third component is selected from the group consisting of pentafluoroethane, 1,1,2,2-tetrafluoroethane, 1,1,1,2-tetrafluoroethane, 1,1,1,2,3,3,3-heptafluoropropane,  
5 1,1,1,2,2,3,3-heptafluoropropane, and 1,1,1,2,2-pentafluoropropane.

7. The refrigerant compositions of claim 1 wherein said first component is difluoromethane.

10 8. The refrigerant compositions of claim 1 wherein said first component is difluoromethane, said second component is trifluoromethane, and said third component is 1,1,1,2-tetrafluoroethane.

15 9. The refrigerant compositions of claim 1 wherein said first component is propane.

20 10. A method for producing refrigeration which comprises condensing said refrigerant compositions of claim 1 and thereafter evaporating said refrigerant compositions in the vicinity of the body to be cooled.

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# INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 92/01160

## I. CLASSIFICATION OF SUBJECT MATTER (If several classification symbols apply, indicate all)<sup>6</sup>

According to International Patent Classification (IPC) or to both National Classification and IPC

Int.Cl. 5 C09K5/04

## II. FIELDS SEARCHED

### Minimum Documentation Searched<sup>7</sup>

Classification System	Classification Symbols
Int.Cl. 5	C09K

Documentation Searched other than Minimum Documentation  
to the Extent that such Documents are Included in the Fields Searched<sup>8</sup>

## III. DOCUMENTS CONSIDERED TO BE RELEVANT<sup>9</sup>

Category <sup>10</sup>	Citation of Document, <sup>11</sup> with indication, where appropriate, of the relevant passages <sup>12</sup>	Relevant to Claim No. <sup>13</sup>
P,A	EP,A,0 430 171 (MATSUSHITA) 5 June 1991 see abstract; claims 1,3 ---	1-8
P,A	EP,A,0 430 170 (MATSUSHITA) 5 June 1991 see abstract; claims 1,3 ---	1-8
P,A	EP,A,0 430 169 (MATSUSHITA) 5 June 1991 see page 1; claims 1-5 ---	1-8
P,A	WORLD PATENTS INDEX LATEST Derwent Publications Ltd., London, GB; AN 91-262361 & JP,A,3 170 591 (MATSUSHITA) 24 July 1991 see abstract --- -/--	1-8

<sup>10</sup> Special categories of cited documents:

"A" document defining the general state of the art which is not considered to be of particular relevance

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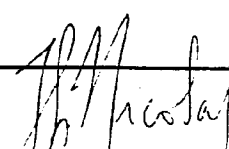
"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.

"&" document member of the same patent family

## IV. CERTIFICATION

Date of the Actual Completion of the International Search	Date of Mailing of this International Search Report
10 JULY 1992	16. 07. 92
International Searching Authority EUROPEAN PATENT OFFICE	Signature of Authorized Officer NICOLAS H.J.F. 



**ANNEX TO THE INTERNATIONAL SEARCH REPORT  
ON INTERNATIONAL PATENT APPLICATION NO. US 9201160  
SA 58505**

This annex lists the patent family members relating to the patent documents cited in the above-mentioned international search report.  
The members are as contained in the European Patent Office EDP file on  
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information. 10/07/92

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP-A-0430171	05-06-91	JP-A- 3172384	25-07-91
		JP-A- 3172386	25-07-91
EP-A-0430170	05-06-91	JP-A- 3170584	24-07-91
		JP-A- 3170587	24-07-91
EP-A-0430169	05-06-91	JP-A- 3170585	24-07-91
		JP-A- 3172385	25-07-91